

Find the mass of 0.89 mol of CaCl_2 .

Determine the number of atoms that are in 0.58 mol of Se.

A bottle of PbSO_4 contains 158.1 g of the compound. How many moles of PbSO_4 are in the bottle?

How many moles of barium nitrate (BaNO_3) contain 6.80×10^{24} formula units?

Find the mass of 1.112 mol of HF.

Determine the number of atoms that are in 1.25 mol of O_2 .

Determine the number of moles of C_5H_{12} that are in 362.8 g of the compound.

How many moles of magnesium bromide (MgBr_2) contain 5.38×10^{24} formula units?

How many formula units are in 1.4 g of PbCl_2 ?

If you burned 6.10×10^{24} molecules of ethane (C_2H_6), what mass of ethane did you burn?

Determine the mass of 2.94×10^{24} molecules of decane ($\text{C}_{10}\text{H}_{22}$).

How many formula units are in 5.1 g of TiO_2 ?

How many formula units are in 5.6 g of H_2S ?

What is the mass of 3.62×10^{24} molecules of methanol (CH_3OH)?

A bottle of KMnO_4 contains 66.38 g of the compound. How many moles of KMnO_4 does it contain?

How many formula units are in 3.5 g of NaOH?